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| **TEST I TREMUJORIT TË TRETË**  **Lënda : Kimi 8**  **Kreu : V**  **Emër Mbiemër \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**   1. **Plotësoni** fjalitë me fjalët që mungojnë: **( 5 pikё)**  * Substancat nistore tё reaksionit quhen\_\_\_\_\_\_\_\_\_\_\_\_\_\_ dhe substancat qё përftohen quhen \_\_\_\_\_\_\_\_\_\_ * Ligji i ruajtjes sё masës thotё: Masa e\_\_\_\_\_\_\_\_\_\_\_\_\_\_ e \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ ёshtё e barabartë me \_\_\_\_\_\_\_\_\_\_ e përgjithshme tё \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. * Një reaksion kimik dallohet nga: ndryshimi i \_\_\_\_\_\_\_\_\_\_\_\_\_\_, çlirimi i \_\_\_\_\_\_\_\_\_\_\_\_\_\_, ndryshimi i \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_, ndryshimi i vlerёs sё \_\_\_\_\_\_, formimi i njё \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. * Procesi i mbrojtjes sё hekurit nga \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ quhet \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.  1. Nga bashkëveprimi i **hidroksidit tё natriumit** me **acid karbonik** përftohet **karbonat natriumi** dhe **ujё.** 2. Shkruani me anё tё formulave kimike skemën e reaksionit kimik. (3) 3. Barazojeni atё. (3) 4. Përcaktoni tipin e reaksionit. (1) 5. Përcaktoni reaktantёt dhe produktet. (2) **(9 pikë)** 6. **Përcakto** shndërrimet e mëposhtme nëse janë fizike **(F)** apo kimike **(K)? (4 pikё)**  * Djegia e karburanteve **( )** * Avullimi i alkoolit **( )** * Fermentimi i rrushit **( )** * Ndryshkja e hekurit **( )**  1. Shkruani **barazimet** e reaksioneve pёr kalimet e mëposhtme: **(4 pikё)**   Ca → CaO → Ca(OH)2 → CaCO3 (grupi A)  Al → Al2O3 → AlCl3 → Al(OH)3 → Al2O3 (grupi B)   1. Përcaktoni **llojin** e reaksionit: **(5 pikё)**   K + H2O → KOH + H2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  AgNO3 + NaCl → ↓AgCl + NaNO3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  P4 + O2 → P2O5 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  CaCO3 → CaO + CO2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  H2CO3 + NaOH → NaHCO3 + H2O \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_   1. **Klasifikoni** substancat e mëposhtme nё: **okside, acide, baza, kripëra: (8 pikё)**   **Emërtoji ato.**  NaCl \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ SO2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  H2CO3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Mg(OH)2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  KOH \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ H2SO3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  Na2O \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ KNO3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_   1. **Klasifikoni** nё okside **bazike** dhe okside **acide**: **(2 pikё)**   CO2\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Al2O3\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  FeO \_\_\_\_\_\_\_\_\_\_\_\_\_ Cl2O5 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_   1. Shkruani 3 rrugë tё ndryshme tё përftimit tё kripës NaCl**. (3 pikё)**   \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ |
| * **Vlerësimi i nxënësve**   Nxënësit vlerësohen sipas pikëzimit tё testit. |

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| **Nota** | **4** | **5** | **6** | **7** | **8** | **9** | **10** |
| **Pikёt** | 0-9 | 10-15 | 16-20 | 21-25 | 26-30 | 31-35 | 36-40 |